

Course ID: HOA241	Course name: INORGANIC CHEMISTRY II					
Cycle: FIRST	Year: SECOND		Semester: IV	ECTS credits: 6		
Course status: MANDAT		DRY	Total course hours Lectures: 45 Laboratory: 45	: 90		
Teaching participants:		Teachers and associates with expertise in the field of Inorganic Chemistry				
Prerequisite for enrollment:		-				
Course aims:		Study of the elements thr of elements compounds.	chemistry of transit cough general, physic and the most in	tion and internal transition cal and chemical properties nportant classes of their		
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Learning outcomes	:	 After the course the student will be able to: explain the similarities and differences of transition and internal transition elements and their compounds with elements and compounds of s- and p-blocks use vertical, horizontal and diagonal similarity indicators to predict the physical and chemical properties of elements and their compounds within d- and f-blocks understand and predict the structures of elemental substances and molecular and crystalline binary and ternary compounds of d-elements 				

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 understand, explain and compare general, physical and chemical properties and methods of obtaining elemental substances, binary and ternary compounds of d- and f-block elements with emphasis on oxides (chalcogenides), halides, hydroxides, oxosalts and oxoacids predict and compare acid-base properties of elemental substances and compounds of d- and f-blocks by groups and periods predict and compare redox properties of elemental substances and compounds of d- and f-blocks by groups and periods predict and compare redox properties of elemental substances and compounds of d- and f-blocks by groups and periods know, explain, predict and distinguish the chemical reactivity of d- and f-block elements and their ionic and molecular compounds within groups and by periods, and based on that predict the products of chemical reactions Teaching methodology: Auditory lectures, laboratory exercises Assessment methods and grading system ¹ : Assessment methods and grading system ¹ : Class attendance S ores and grading Calss attendance S ores Class attendance S ores S ores						
Teaching methodology:Auditory lectures, laboratory exercisesGrading criteriaGrading criteriaMaximal scoreRequired score1.Class attendance2.Class activities3.Midterms2 × 252 × 144.Final exam*4022Total100Scores and gradingScores and gradingScores and gradingScoreGradeGradeScore(BiH)(ECTS)< 555< 555< 55-646< 65-747 < 75-848 < 95-10010A*oral exam after student successfully completes midtermsMandatory literature:1.Atkins, P., & Overton, T. (2010). Shriver and Atkins' inorganic chemistry. Oxford University Press USA		 understand, explain and compare general, physical and chemical properties and methods of obtaining elemental substances, binary and ternary compounds of d- and f-block elements with emphasis on oxides (chalcogenides), halides, hydroxides, oxosalts and oxoacids predict and compare acid-base properties of elemental substances and compounds of d- and f-blocks by groups and periods predict and compare redox properties of elemental substances and compounds of d- and f-blocks by groups and periods predict and compare redox properties of elemental substances and compounds of d- and f-blocks by groups and periods predict and periods predict and compare redox properties of elemental substances and compounds of d- and f-blocks by groups and periods know, explain, predict and distinguish the chemical reactivity of d- and f-block elements and their ionic and molecular compounds within groups and by periods, and based on that predict the products of abamical reactions 				
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morganie enemistry: omora omversity rress, oom		inorganic chemistry.	Oxford University Pres	ss, USA.		

¹The grading structure for each subject is determined by the Council of the organizational unit before the beginning of the academic year in which the subject is taught as per Article 64, paragraph 6 of the Law on Higher Education of Sarajevo Canton

²The Senate of the higher education institution, as an institution, or the Council of the organizational unit of the higher education institution, as a public institution, determines by a special decision, which is published on its website before the beginning of the academic year obligatory, mandatory and recommended textbooks and manuals, as well as other recommended literature based on which exams are prepared and taken as per Article 56, paragraph 3 of the Law on Higher Education of the Sarajevo Canton

Form SP2

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2. Emira Kahrović, Nevzeta Ljubijankić, Praktikum anorganske hemije, Prirodno-matematički fakultet, 2011.	
3. Filipović, D. I., Lipanović, D. S. (1995). Opća i anorganska	
kemija. II. dio. Kemijski elementi, njihove elementarne	
tvari i spojevi, Školska knjiga, Zagreb.	
4. Greenwood, N. N., & Earnshaw, A. (2012). Chemistry of	
the Elements. Elsevier.	
Supplementary literature:	
1. Miessler, G. L. And Tarr, D. A. (1999). Inorganic Chemistry,	
Prentice-Hall.	